

A 51  
Infrarot - Absorptions - Spektroskopie  
22.2.17



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Praktikum: A51

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Betreuer: *[Signature]*

## Objective

The objective of the experiment is to measure the rotational-vibrational spectrum of carbon monoxide CO with a Fourier Transform infrared spectrometer FTIR.

Afterwards we interpret the spectrum and determine the resonant vibrational frequencies, rotational constants, moments of inertia and bond length of CO.

## Basics

The FTIR provides ~~an~~ an absorption spectrum by using a Michelson interferometer. All wavelengths are measured simultaneously.

The obtained interferogram is converted into an absorption spectrum using a fast Fourier transformation - algorithm.

Absorption occurs by changing the vibrational and rotational states of the sample, but only certain energies can be absorbed. This leads to the individual peaks in the spectrum.

The oscillation in the CO-Molecule can be <sup>approximately</sup> described as the oscillation of point masses in a harmonic potential.

Related Schrödinger equation:  $-\frac{\hbar^2}{2\mu} \frac{d^2}{dx^2} \psi(x) + \frac{1}{2} kx^2 \psi(x) = E \psi(x)$

Solved by  $E_{vib}(v) = hc \tilde{\nu}_0 (v + \frac{1}{2})$   $v = 0, 1, 2, \dots$   $\tilde{\nu}_0 = \frac{1}{2\pi c} \sqrt{\frac{k}{\mu}}$

The rotation of the molecule follows the principle of the rigid rotator.

Energy states  $E_{rot}(J) = hcB J(J+1)$  ;  $J = 0, 1, 2, \dots$   $B = \frac{h}{8\pi^2 c I} = \frac{\hbar^2}{4\pi c I}$

B is ~~the~~ the rotational constant that is used to derive the bond length and the moment of inertia.

$$I = \frac{\hbar^2}{4\pi c B} ; I = \mu r^2$$

In order to interpret the spectrum following selection rules must be observed:

Source: ~~experiment~~ Description of experiment  
Athens, Physikalische Chemie, WILEY-VCH, Weinheim 2013

global rules: for rotational ~~excitation~~ <sup>excitation</sup> a molecule must have a permanent electric dipole moment.  
 for vibrational excitation the electric dipole moment has to change as a function of the nuclei distance.

Also the specific selection rules limit the possible quantum numbers.

Two atomic oscillator  $\Delta v = \pm 1$  ; ~~two~~ <sup>two</sup> atomic rigid rotator  $\Delta J = \pm 1$

If one compares the reality with the previous considerations, one can see deviations. These come from the anharmonicity of the ~~oscillator~~ oscillator. Therefore corrections have to be implied.

Equations for data evaluation

Fundamental:  $\tilde{\nu}_R(j) - \tilde{\nu}_P(j) = B_1(4j+2)$  ;  $\frac{1}{2}(\tilde{\nu}_R(j-1) + \tilde{\nu}_P(j)) = \tilde{\nu}_{00} - \alpha_e j^2$   
 $\tilde{\nu}_R(j-1) - \tilde{\nu}_P(j+1) = B_0(4j+2)$

Overtone:  $\tilde{\nu}_R(j) - \tilde{\nu}_P(j) = B_2(4j+2)$  ;  $\frac{1}{2}(\tilde{\nu}_R(j-1) + \tilde{\nu}_P(j)) = \tilde{\nu}_{200} - 2\alpha_e j^2$   
 $\tilde{\nu}_R(j-1) - \tilde{\nu}_P(j+1) = B_0(4j+2)$

Rotational constant  $B_v = B_e - \alpha_e(v + \frac{1}{2})$

Anharmonicity  $\frac{\tilde{\nu}_{v+1,0} - \tilde{\nu}_{v,0}}{v} = \tilde{\nu}_e - \tilde{\nu}_e x_e(v+1)$

## Experimental procedure

First a background spectrum was measured, so there was no sample between the light source and the detector. After that another spectrum was measured but this time the CO cuvette is in the beam path. The resolution value amounts to 0.5 and 5 scans per sample were measured.

## Data evaluation

The rotational constants  $B_0, B_1$  and  $B_2$  are determined by the gradient of the linear regression of the measured wavenumbers (graphs attached). With those constants the moments of inertia  $I$  are calculated by the equation given in the basics. Also the bondlength  $R$  of the CO-Molecule is calculated that way:  $R = \sqrt{\frac{I}{\mu}}$  (conversion of the equation in the basics). The calculation of the anharmonicity constant  $x_e$  is a consequence of the linear regression of the equation  $\frac{\tilde{\nu}_{v+1}}{v} = \tilde{\nu}_e - \tilde{\nu}_e x_e (v+1)$ . The gradient is given by the Program originlab and to get  $x_e$  it had to be divided by  $\tilde{\nu}_e$ . All results are listed in a table in the annex.

## Critical discussion

$\tilde{\nu}_{0 \rightarrow 1}$  exists more often than  $\tilde{\nu}_{0 \rightarrow 2}$  which is a result of the Boltzmann-distribution so the intensity of the harmonic oscillator is smaller.

As the results of the rotational constants deviate from the literature all the other results cannot be exactly equal with the literature, too.

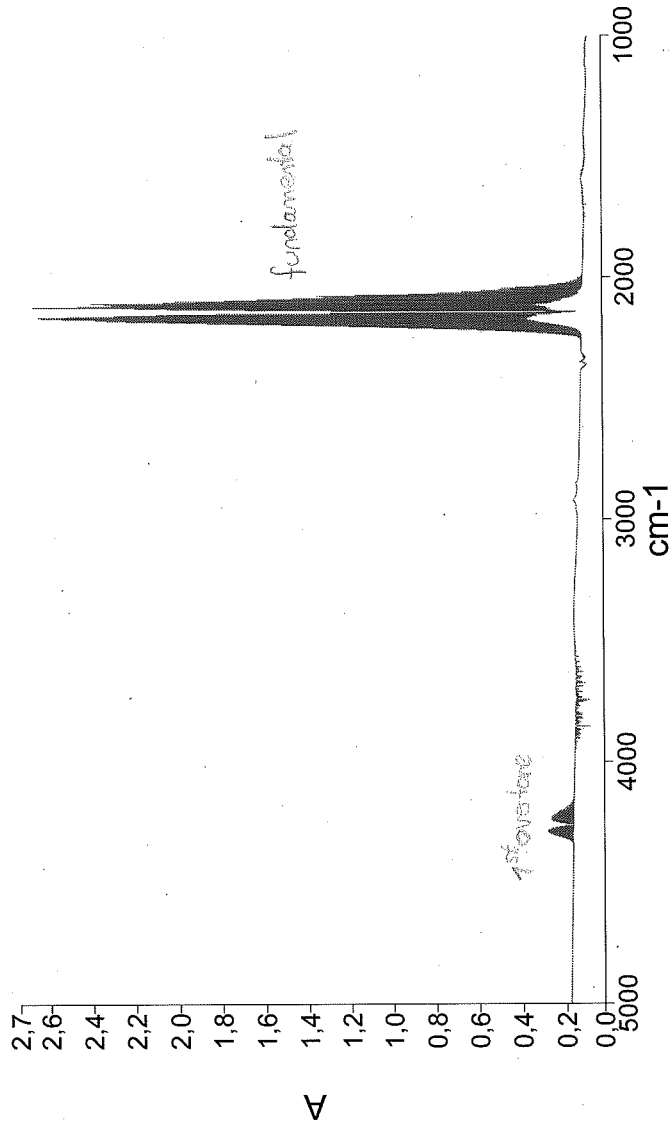
Centrifugal forces stretch the bond of the molecule, so the rotational constants get smaller and cause asymmetrical behaviour.

The Q-branch shouldn't be in the CO-spectrum because of the selection rules. In the fundamental there is a small peak in the Q-branch which indicates that the carbon monoxide molecule is not an ideal harmonic oscillator. Only in this model the selection rules forbid the Q-branch.

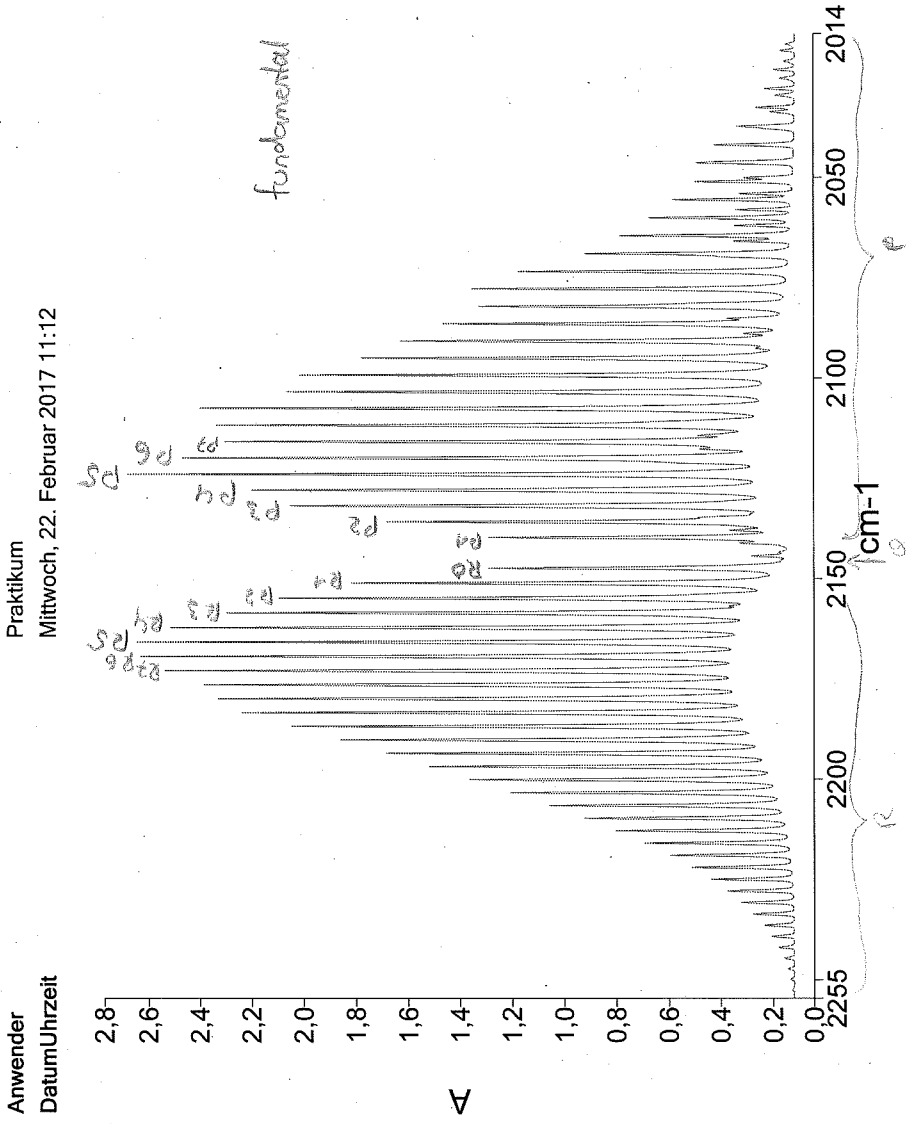
## Source for literature data:

- description of experiment
- Atkins, Physikalische Chemie, WILEY-VCH, Weinheim 2013
- Pyler, E.K.; Blaine, L.R.; Connor, W.S. Opt. Soc. Amer. 1955, 45, 102-106
- Rank, D.H.; Guenther, A.H.; Sakson, G.D.; Shearer, I.A. J. Opt. Soc. Amer. 1957, 47, 688-689

Anwender  
Datum/Uhrzeit  
Praktikum  
Mittwoch, 22. Februar 2017 11:01



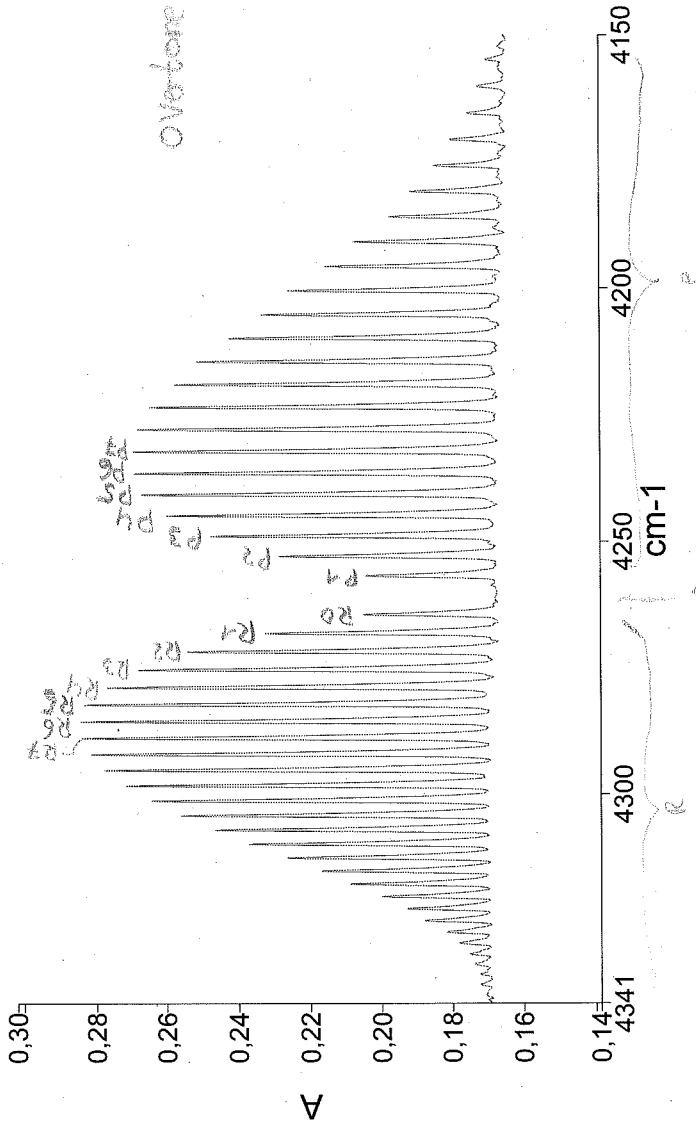
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Praktikum47_2	Sample 016 By Praktikum date Mittwoch, Februar 22 2017



Sample 016 By  
Praktikum date  
Mittwoch, Februar 22  
2017

Praktikum47\_2

Anwender: Praktikum  
Datum/ Uhrzeit: Mittwoch, 22. Februar 2017 11:03



Probenname	Beschreibung
Praktikum47_2	Sample 016 By Praktikum date Mittwoch, Februar 22 2017



Quellspektren- Ergebnisse
Name des Spektrums

Liste der Peakflächen/Peakhöhen		
Name des Peaks	X (cm-1)	Y (A)
1	4312,53	0,2265
2	4309,82	0,2372
3	4307,03	0,2465
4	4304,19	0,2559
5	4301,27	0,2641
6	4298,27	0,2718
7	4295,19	0,2777
8	4292,06	0,2813
9	4288,85	0,2842
10	4285,57	0,2842
11	4282,21	0,2833
12	4278,79	0,2767
13	4275,3	0,268
14	4271,74	0,254
15	4268,1	0,2324
16	4252,86	0,2285
17	4248,88	0,2474
18	4244,82	0,26
19	4240,69	0,267
20	4236,5	0,269
21	4232,23	0,2691
22	4227,9	0,2681
23	4223,5	0,2645
24	4219,04	0,2572
25	4214,5	0,2511
26	4209,89	0,2421

R10  
R9  
R8  
R7  
R6  
R5  
R4  
R3  
R2  
R1  
R0  
P1  
P2  
P3  
P4  
P5  
P6  
P7  
P8  
P9  
P0

Liste der Peakflächen/Peakhöhen		
Name des Peaks	X (cm-1)	Y (A)
27	4205,22	0,2331
28	4200,49	0,2257
29	3842,52	0,1577
30	3417,13	0,1626
31	2241,94	0,1782
32	2239,2	0,2045
33	2236,43	0,2368
34	2233,62	0,2755
35	2230,77	0,3201
36	2227,89	0,3735
37	2224,96	0,4365
38	2221,99	0,511
39	2218,98	0,596
40	2215,93	0,6923
41	2212,85	0,8039
42	2209,72	0,9239
43	2206,57	1,0588
44	2203,37	1,2064
45	2200,13	1,3632
46	2196,86	1,5207
47	2193,54	1,6836
48	2190,2	1,8613
49	2186,81	2,0503
50	2183,4	2,2362
51	2179,93	2,3282
52	2176,45	2,3835
53	2172,91	2,5383
54	2169,38	2,6341
55	2165,79	2,6491
56	2162,15	2,5173

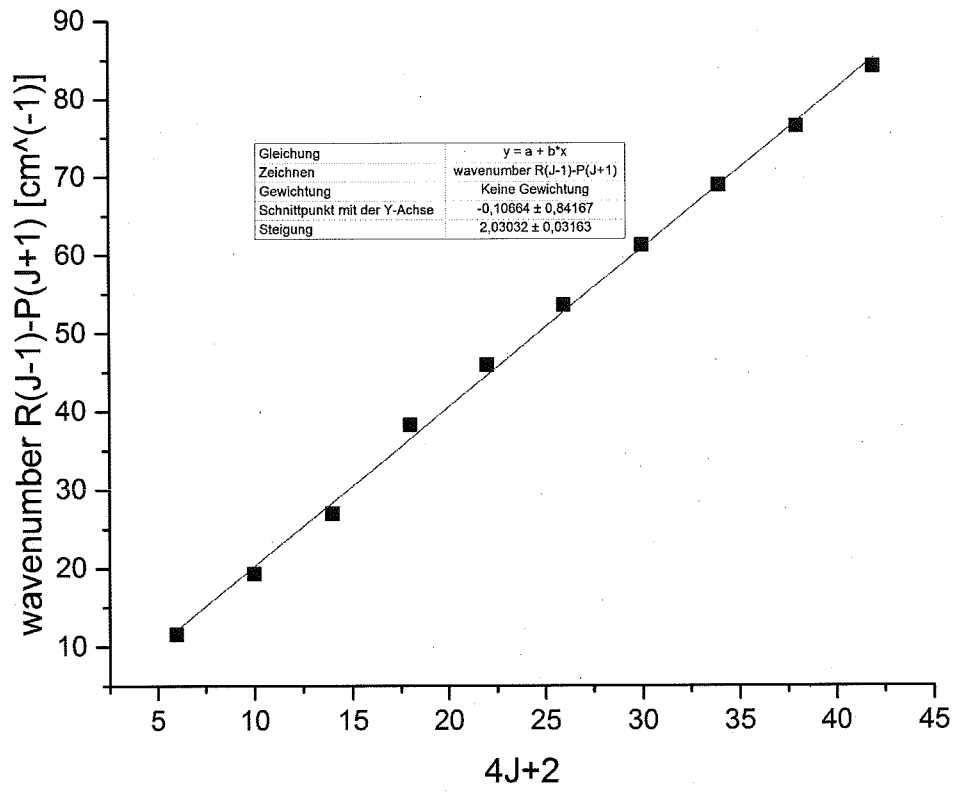
R10  
R5  
R8  
R7  
R6  
R5  
R4  
R3

Liste der Peakflächen/Peakhöhen		
Name des Peaks	X (cm-1)	Y (A)
58	2154,79	2,0964
59	2151,01	1,8173
60	2147,3	1,2876
61	2144,28	0,2791
62	2141,03	0,3248
63	2139,63	1,289
64	2137,82	0,3632
65	2135,73	1,6825
66	2131,77	2,0511
67	2127,85	2,2008
68	2123,9	2,6783
69	2119,87	2,4705
70	2115,8	2,303
71	2114,22	0,4882
72	2111,71	2,3388
73	2107,55	2,4003
74	2103,44	2,0664
75	2099,28	2,0147
76	2095,05	1,7803
77	2090,8	1,6299
78	2088,91	0,3084
79	2086,51	1,4679
80	2082,2	1,326
81	2077,8	1,3516
82	2073,5	1,1763
83	2069,08	0,9196
84	2065,99	0,3476
85	2064,61	0,7855
86	2062,07	0,3439
87	2060,13	0,6739

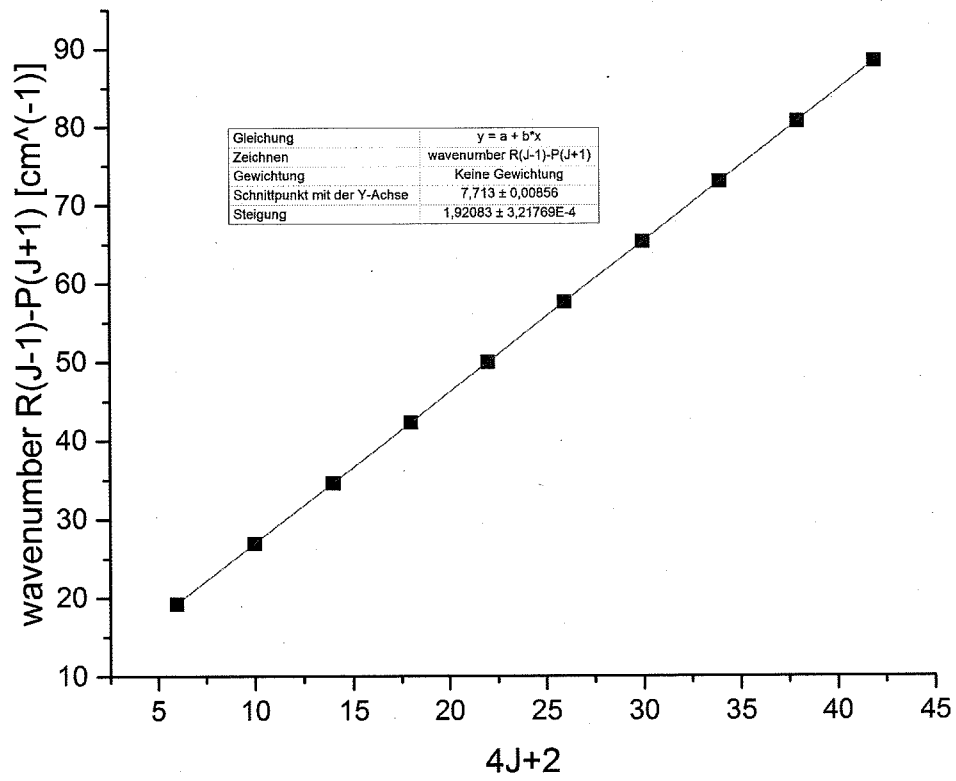
R2  
R1  
R0  
Q  
Q  
P1  
  
P2  
P3  
P4  
P5  
P6  
P7  
  
P8  
P9  
P10

Liste der Peakflächen/Peakhöhen		
Name des Peaks	X (cm-1)	Y (A)
89	2055,61	0,582
90	2054,13	0,3242
91	2051,08	0,4994
92	2050,12	0,3104
93	2046,41	0,4916
94	2041,91	0,4222
95	2037,29	0,3326
96	2033,67	0,2073
97	2032,58	0,2658
98	2029,49	0,1883
99	2027,88	0,2276
100	2025,27	0,1718
101	2023,14	0,1931
102	2018,38	0,1691

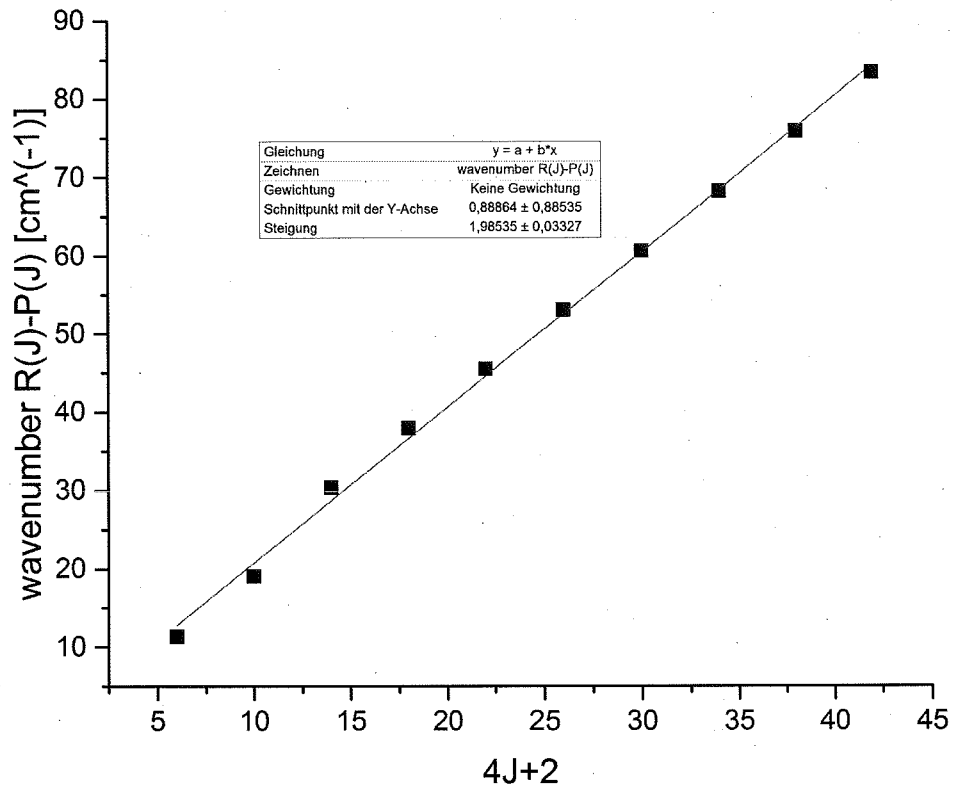
### B\_0 fundamental



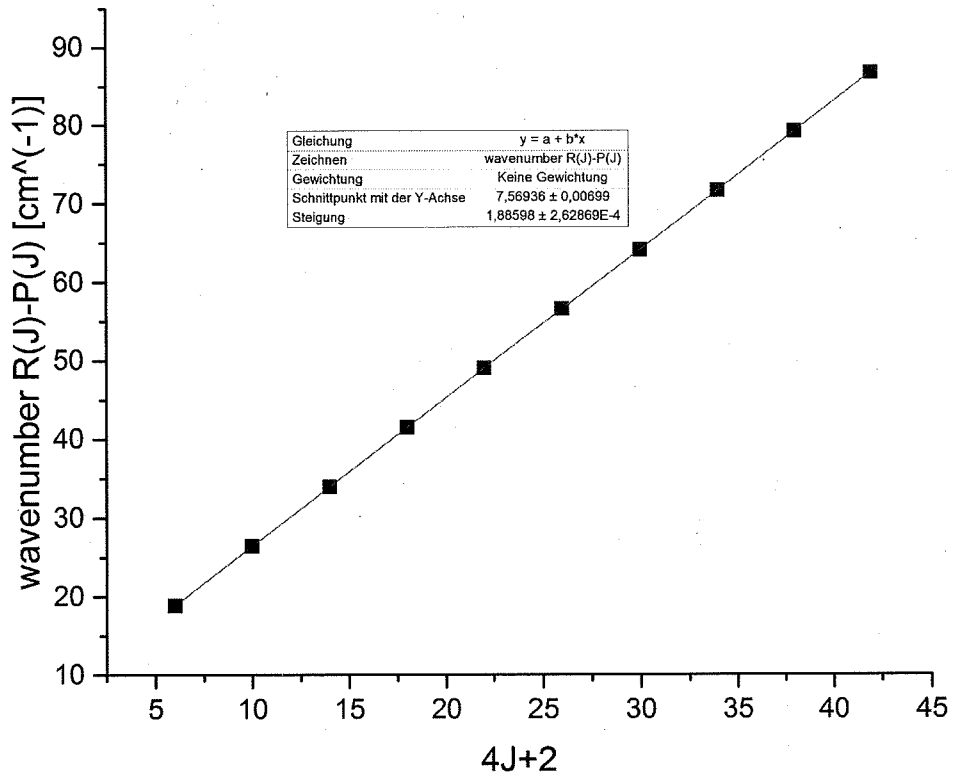
### B\_0 overtone



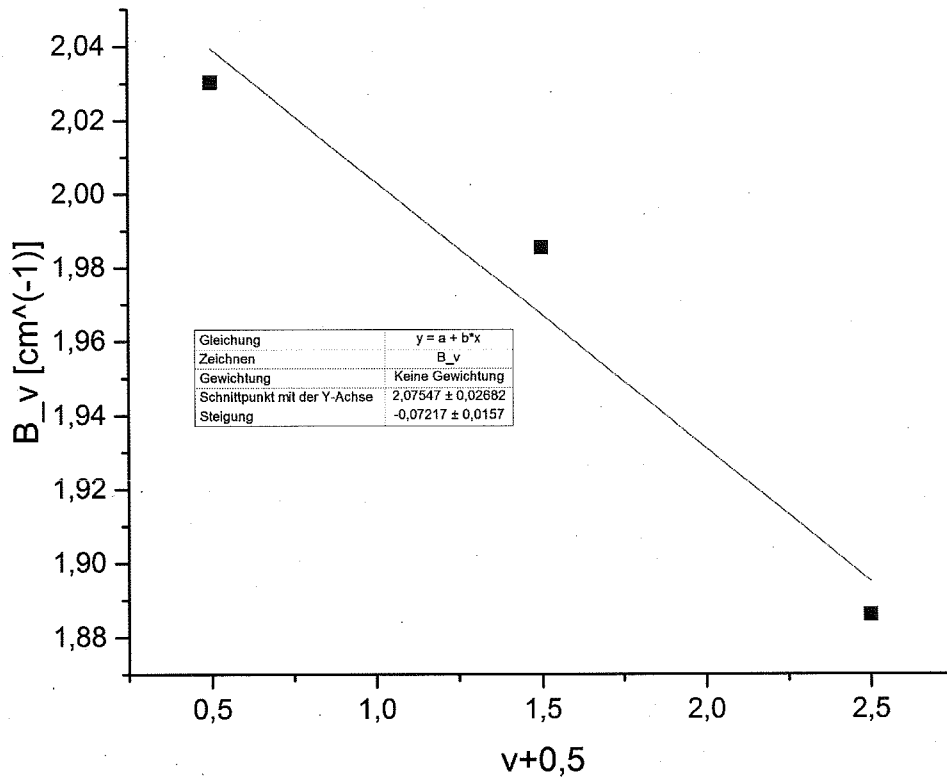
### B1 fundamental



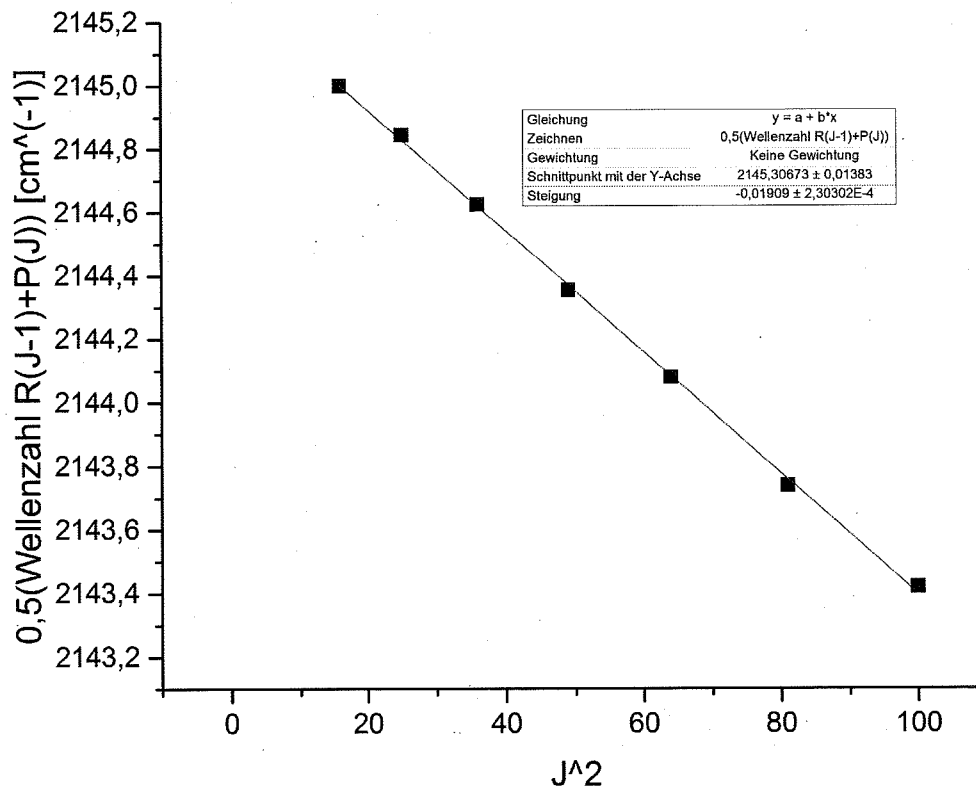
### B2 overtone



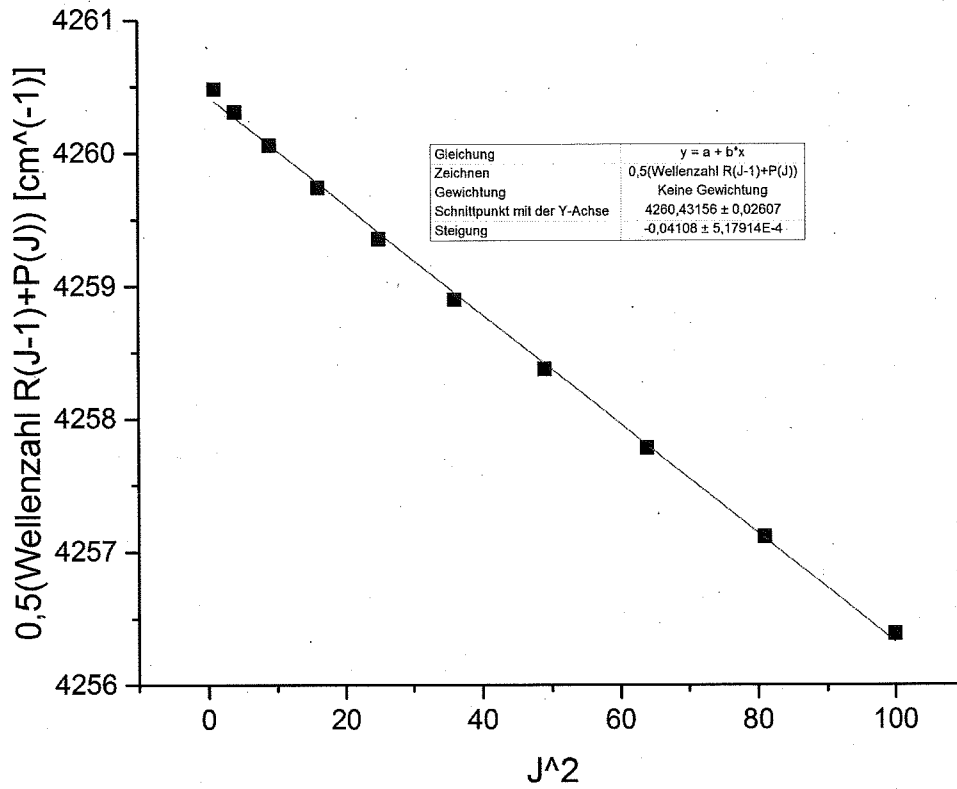
B\_v



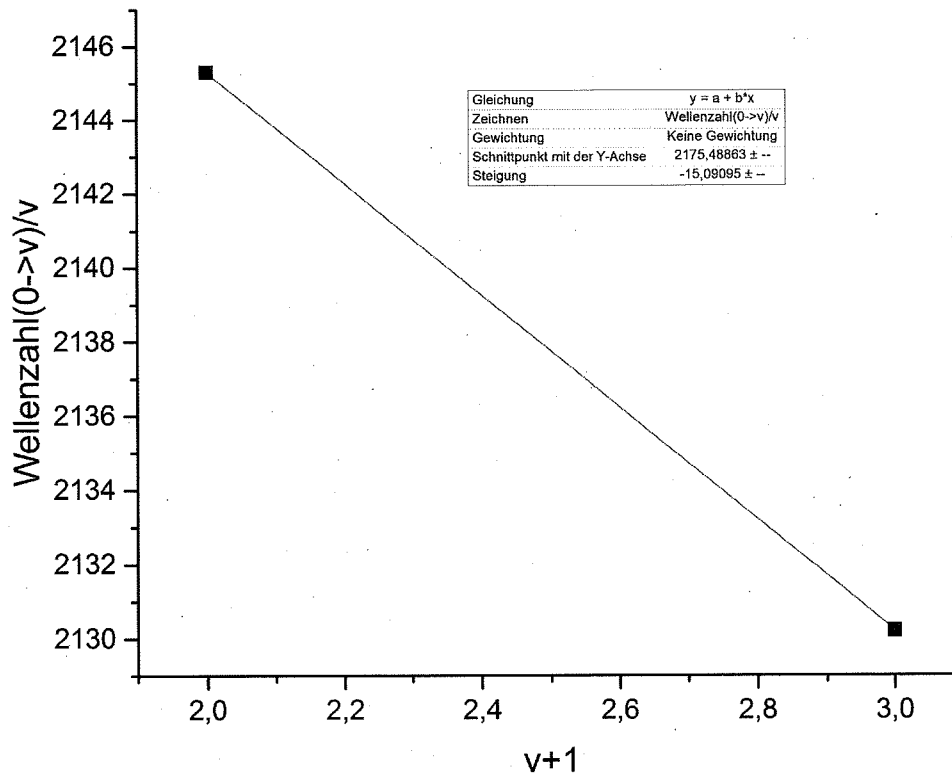
J^2 fundamental



### J<sup>2</sup> overtone



### anharmonicity





measurand	measured value	error	literature	
B <sub>0</sub>	2,03032	0,01597588	1,9326 cm <sup>-1</sup>	} → graphics named like measurand
B <sub>1</sub>	1,98535	0,03327	1,9051 cm <sup>-1</sup>	
B <sub>2</sub>	1,88598	0,00026287	1,8875 cm <sup>-1</sup>	
I <sub>0</sub>	1,37236E-46	1,0799E-50	1,456E-46 kg*m <sup>2</sup>	
I <sub>1</sub>	1,40345E-46	2,3519E-50	1,4693E-46 kg*m <sup>2</sup>	
I <sub>2</sub>	1,47739E-46	2,0592E-52	1,4830E-46 kg*m <sup>2</sup>	
R <sub>0</sub>	<u>7,52794E-10</u>	7,1855E-42	1,1306E-10 m	
R <sub>1</sub>	<u>7,61272E-10</u>	1,5475E-41	1,1358E-10 m	
R <sub>2</sub>	<u>7,8107E-10</u>	1,3206E-43	1,1411E-10 m	
$\tilde{V}_{(0 \rightarrow 1)}$	2145,30673		2143,2 cm <sup>-1</sup>	→ graphic "j <sup>2</sup> fundamental"
$\tilde{V}_{(0 \rightarrow 2)}$	2130,21578		4259,8 cm <sup>-1</sup>	→ graphic "j <sup>2</sup> overtone"
$\tilde{V}_e$	2175,48863		2169,8 cm <sup>-1</sup>	→ graphic "anharmonicity"
x <sub>e</sub>	0,006936809		6,125E-03 cm <sup>-1</sup>	
α <sub>e</sub> (fundamental)	0,01909		1,75E-2 cm <sup>-1</sup>	
α <sub>e</sub> (overtone)	0,02054			
B <sub>e</sub>	2,07547	0,02682	1,8313 cm <sup>-1</sup>	→ in the graphic called "B <sub>v</sub> "

For the errors in the table above the Gaussian law of errors was used:  $\Delta y = \sqrt{\sum_i \left(\frac{\partial y}{\partial x_i} \Delta x_i\right)^2}$

$$\Rightarrow \Delta I_i = \frac{h}{8\pi^2 c \cdot B_i^2} \Delta B_i$$

$$\Rightarrow \Delta R_i = \frac{1}{2} \cdot \frac{1}{\sqrt{I_i}} \cdot \Delta I_i$$

The error of B<sub>i</sub> was given by the statistical error of the linear regression with the program OriginLab (shown in the graphs attached)

Source of literature:

1. Pyle, E.K.; Blaine, L.R.; Connor, W. S. J. Opt. Soc. Amer. 1955, 45, 702-706
2. Renk, D.H.; Guenther, A.H.; Salsera, G.D.; Shearer, T. A. J. Opt. Soc. Amer. 1957, 47, 686-688